

“Easy Does It”: More Molarity Practice

1. How many grams of K_2O are needed to prepare a 200.0mL 0.25M solution?
2. How many grams of $MgCl_2$ are needed to prepare a 150.mL 0.50M solution?
3. What is the concentration of a solution made from mixing 25g of $C_6H_{12}O_6$ and diluting with water to 250 mL?
4. What is the concentration of a solution prepared by mixing 3.2g of $NaOH$ and diluting to 0.50 L with water?
5. How many milliliters are required to prepare a 0.45M solution of $Cu(OH)_2$ if 2.3g is used?
6. How many liters are required to prepare a 0.60M solution of CuO if 5.72g is used?
7. How would you prepare a 500.0mL 0.200M solution of $AgNO_3$ if you have a 0.500M solution to use?
8. How would you prepare a 200.0mL 0.200M solution of $AgNO_3$ if you have a 0.500M solution to use?
9. How many milliliters of a 2.00M solution of $NaOH$ would you need to prepare 500.0mL of a 1.00M solution?
10. $AgNO_3 + CaCl_2 \rightarrow Ca(NO_3)_2 + AgCl$
If 250. mL of 0.500M $AgNO_3$ are added to $CaCl_2$, how many grams of $AgCl$ will be produced?
11. $CsBr + Al(OH)_3 \rightarrow AlBr_3 + CsOH$
 - a. If 25 g of $AlBr_3$ was produced when 125 mL of a $CsBr$ solution was added, what was the molarity of the $CsBr$ solution?
 - b. How many mL of 0.50 M $CsBr$ are required to produce 100 g of $CsOH$?